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## WRITING BALANCING CHEMI...

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### WORKSHEET - BALANCING CHEMICAL EQUATIONS

Worksheet: Writing and Balancing Chemical Reactions  
 1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.  
 a.  $\text{Cu (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{CuO (s)}$   
 b.  $\text{H}_2\text{O (l)} \rightarrow \text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$   
 c.  $\text{Fe (s)} + \text{H}_2\text{O (g)} \rightarrow \text{H}_2 \text{ (g)} + \text{Fe}_3\text{O}_4 \text{ (s)}$   
 d.  $\text{AsCl}_3 \text{ (aq)} + \text{H}_2\text{S (aq)} \rightarrow \text{As}_2\text{S}_3 \text{ (s)} + \text{HCl (aq)}$   
 e.  $\text{CuSO}_4 \cdot 5 \text{ H}_2\text{O (s)} \rightarrow \text{CuSO}_4 \text{ (s)} + \text{H}_2\text{O (l)}$

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Balancing Chemical Equations Answer Key  
 Balance the equations below:  
 1)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$   
 2)  $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$   
 3)  $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$   
 4)  $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$   
 5)  $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$   
 6)  $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$   
 7)  $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$   
 8)  $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$   
 9)  $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$   
 10)  $1 \text{ FeCl}_3 + 3 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_3 + 3 \text{ NaCl}$   
 11)  $4 \text{ P} + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$   
 12)  $2 \text{ Na} + \dots$

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### BALANCING CHEMICAL EQUATIONS

BALANCING CHEMICAL EQUATIONS  
 The Law of Conservation of Mass says that in a chemical reaction, matter is neither created nor destroyed. Practically speaking, whatever atoms begin in the reactants of a chemical reaction, the same number of atoms of each element must appear in the products. The process of balancing chemical equations is intended to reflect reality, that is, to give us a true picture of...

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